**MWEMBEZHI SECONDARY SCHOOL**

**CHEMISTRY LESSON PLAN**

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| --- | --- | --- | --- | --- |
| **GRADE**  |  **TOPIC** |  **SUB TOPIC** | **DURRATION** | **DATE** |
| 11 | Acids, bases and salts | acids | \_\_\_ min | \_\_\_/\_\_\_/\_\_\_ |

**RATIONALE**: Strong acids are acids that completely ionise in water. This lesson will help learners to identify strong acids. This is the second lesson of three in the series. Group work will be used to execute the lesson.

**SPECIFIC OUTCOMES, LSBAT: -** Identifying strong acids

 -give the characteristics of acids

 - Describe the meaning of weak, strong, dilute and

 Concentrated acid.

**PRE-REQUISITE SKILLS:** learners have knowledge on the chemical formula

**TEACHING AND LEARNING AIDS:** chemistry key pointspupils book 11, reals acids (sulphuric acid)

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**REFFERENCES :**

 **LESSON PROGRESSION**

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| **STAGE** | **T/L ACTIVITY** | **LEARNING POINTS** |
| **INTROD-CTION** | Reflect on the previous lesson by ask questions |  |
|  | H2SO4(aq) 2H3O+(aq) + SO42-(aq)HCL(aq) H3O+(aq) + Cl-(aq) What does the single arrow mean  **ACTIVITY 2**  What are weak acidsAnd they partially ionize as follows H3PO4(aq) H+(aq) + PO43-(aq) H2CO3(aq) 2H+(aq) + CO2-(aq CH3COOH(aq)  H+(aq) +CH3COO-(aq) The double arrows in the above examples means that the reactions are reversible and the acids ionize partially in water. | The single arrow shows that the substance completely ionizes in water.* Weak acids are acids that partially ionize in water

Because of partial ionization some acid particles remain. The remaining un-ionized acid particles remain unchanged and they can recombine to form acid molecule. Examples of weak acids are phosphoric acid, carbonic acid and methanoic acid. |
| **D** **E** **V** **E** **L** **O** **P** **M** **E** **N** **T** |
|  **C** **O** **N** **L** **U** **S** **I** **O** **N** | Discuss the following points;* Strength as degree of ionization,
* Concentration as the number of ions per volume of solution.

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EVALUATION: ………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………….